

# Physical properties of esters

- Because they don't possess OH groups, esters cannot form H-bonds with other ester molecules. As a result, *esters have lower boiling points than carboxylic acids and alcohols that have approximately the same molar mass.*
- They do have a C=O bond (polar) so their boiling points are between alcohols and alkanes.
- Water molecules can H-bond to esters, at the oxygen atoms. This makes *low molecular weight* esters water-soluble.

